

Supporting Documents for Qualitative Metric 2.2.1



Submitted to NAAC

By

**Gandhi Institute For Technology (GIFT),
Bhubaneswar**

2.2.1 Content

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2.2.1 Advanced & Slow learners

Every year the institution organizes Orientation and induction programmed for the new batch students. This programme would help both new students and their parents to get comfortable with the institution, facilities provided, rules and regulation etc.

Bridge classes are conducted in the beginning to lift the students to next level of higher education. The classes are conducted as per the specially designed curriculum with all basic science subjects, Basic English and computer fundamentals. The Syllabus is pre-designed keeping in point of view of psychology and intellectual level of new-comers.

The institution assesses the learning level of learning of the students in two ways at the beginning of the programme. Students enrolled in various disciplines, have to appear an entry level test (Snap Test) conducted by the institution. Based on the performance in the test and then, their +2 marks, they are identified as either advanced (Fast) or slow (weak) learners. In order to motivate them, special tutorial classes are arranged to bridge the gap between the slow learners and advanced learners. Different workshops and induction programme are organized to enhance their skills and make them confident to compete in the main stream.

Strategies adopted for slow learners:

- For the slow learners remedial classes are conducted.
- Simple and standard lecture notes/course materials are provided to the students. These notes are prepared by the subject experts and verified by HOD.
- Faculties who are assigned as mentors counsel the students and monitor their academic performance regularly and interact frequently to understand, assist and support the students towards their better learning.
- Mentors communicate regularly with the parents and also send them SMS along with the progress of the student in regular intervals.
- Extra classes are organized to clarify the doubts. Difficult topics are re-explained and reputedly taught for better understanding of weak students.
- Appropriate counseling with additional attention is done to gradually increase the student's attendance, make them regular in classes and subsequent results.

Strategies adopted for advanced learners:

- High performing students (Fast/Advanced learner) are identified on the basis of internal assessment, previous board marks and active involvement in classroom.
- Advanced learners are provided coaching classes for different competitive exams.
- Advanced learners are encouraged to enroll in NPTEL and TCS ION etc. like advanced on line courses.
- Students are encouraged to take up different micro projects to inculcate research orientation

and practical knowledge for minor and major projects.

- Students are guided by the irrespective branch mentors for various competitive exams like GATE, IES, CAT and NET etc.
- Students are encouraged to participate and present papers in various seminars/conferences/Workshops organized inside the institution and outside colleges.
- Involving students in different consultancy and research projects and to file for own patents.
- They are also encouraged to actively participate in different state and national level competitions like Hackathon and Robotics etc.

Bridge course

To bridge the gap between the students and to provide a common learning platform the institution conducts bridge course, a special programmes after admission. Bridge course for newly admitted students is conducted every year before the commencement of the first semester classes. The main objective of the course is to bridge the gap between subjects studied at Pre-university level and subjects they would be studying in engineering. The syllabus for the course is framed in such a way that equal importance is given to both Engineering discipline and personality development which includes soft skills, sports and cultural Activities. The duration of this course is 25 days with 100 hrs. The students are trained in subjects such as Mathematics, Computer Science and English apart from other co-curricular activities. Syllabus for Mathematics includes concepts learnt at the PU level and also additional concepts which are essential prerequisites for engineering course. The afternoon sessions will be on soft skills, sports and cultural activities. This will help the students to have a smooth transition to the Engineering course. The sound grasp of the fundamentals of the engineering subjects lays the strong foundation for the entire four year programme.



**GANDHI INSTITUTE FOR TECHNOLOGY
BHUBANESWAR**

Date: 18th June, 2021

Ref No:01/06

NOTICE

Bridge Course -2021

It is hereby informed to all the newly inducted 1st year students of B.Tech that the **Bridge Course** for this academic session will commence from 20th June as per the schedule given below.

9.30AM-12.30PM	12.30PM-1.30PM	1.30PM-2.30PM
Theory Classes (Math/Phy/Chem/Eng/Computer/Robotics)	Lunch	Lab Classes (Phy/Chem/Eng/Computer/Robotics)


First Year **Coordinator**


COPY TO:-

All Deans/All HoDs, / Coordinator/ Hostel Wardens/ All College Notice Boards/ All
Hostel Notice Boards



**GANDHI INSTITUTE FOR TECHNOLOGY
BHUBANESWAR**

Ref No:01/05

Date:15th may,2021

NOTICE

Bridge Course - 2021

It is hereby informed to all the newly inducted 1st year students of B.Tech that the **Bridge Course** for this academic session will commence from 21st May as per the schedule given below.

9.30AM to 12.30PM-Theory classes
(Phy/Chem/Math/Computer/Robotics)

12.30PM to 1.30PM-Lunch Break

1.30PM to 3.30PM-Lab classes
(Phy/Chem/Math/Computer/Robotics)

Dev...
First Year Coordinator

COPY TO:-

**All Deans/All HoDs, / Coordinator/ Hostel Wardens/ All College Notice Boards/ All
Hostel Notice Boards**

Time Table [BRIDGE COURSE 2021]

w.e.f: 20-06-2021

Time→ Day↓	9.30-10.30am	10.30-11.30am	11.30-12.30pm	12.30-1.30pm	1.30-2.30pm	2.30-3.30pm	3.30-4.30pm	
Mon	Math P. Sahu	Computer Lab M. Kunar			LUNCH	ECS	Chem	HOBBY
						P. Mishra	Dr. B Mandal	
Tue	Phy Dr. B P Nayak	Math P. Sahu	ECS P. Mishra	English Communication BK Rout		HOBBY		
				Phy Lab B P Nayak/ J. Mohanty		HOBBY		
Wed	ECS P. Mishra	Chem Dr. B Mandal	BCE A Panda/ M. Kunar	Computer Lab M.Kuanr		HOBBY		
				Chem Lab Dr. B Mandal/ P.Pandiya		HOBBY		
Thu	Math P. Sahu	Phy Dr. B P Nayak	Chem Dr. B Mandal	Robotics Lab A Nayak/Amrutashu Panigrahy		HOBBY		
Fri	ECS P. Mishra	Robotics K. N. Hota	Phy Dr. B P Nayak					
Sat	Math P. Sahu	Phy Dr. B P Nayak	Chem Dr. B Mandal					

Subject (Theory)	Teacher(s)
English Communication Skill(EC)	P Mishra
Physics(Phy)	Dr. B P Nayak
Chemistry(Che m)	Dr. B Mandal
Mathematics(Math)	P. Sahu
Basics of Computer Engineering(BC E)	A Panda/ M. Kunar
Robotics	K.N.Hota

HOBBY: Dr. H. B. Panigrahy/ Asit Pahadsingh to look after

Lab Class	Teacher(s)
English Communication	B.K.Rout
Computer Lab	M. Kunar
Phy Lab	B P Nayak/ J. Mohanty/ R.R. Padhi
Chem Lab	H.B. Panigrahy / P Pandya
Robotics	A Nayak/ Amrutanshu Panigrahy

Time Table [BRIDGE COURSE-2021]

w.e.f 21-05-2021

Time→ Day↓	9.30-10.30am	10.30-11.30am	11.30-12.30pm	12.30-1.30pm	1.30-3.30pm	3.30-4.30pm
Mon	ECS	Math	Chem	LUNCH	Computer Lab	HOBBY
	S Tripathy	A K Panigrahi	B Mandal		A Panda	
Tue	Phy	ECS	Math		English Communication	HOBBY
	B P Nayak	S Tripathy	A K Panigrahi		S Tripathy/P Mishra	
Wed	Chem	ECS	BCE		Phy Lab	HOBBY
	B Mandal	S Tripathy	A Panda		B P Nayak/ R R Padhi	
Thu	Math	Phy	Chem		Chem Lab	HOBBY
	A K Panigrahi	B P Nayak	B Mandal		B Mandal/ P Pandya	
Fri	ECS	Robotics	Phy		Computer Lab	HOBBY
	S Tripathy	A Nayak	B P Nayak		A Panda	
Sat	Math	Chem	Phy		Robotics Lab	HOBBY
	A K Panigrahi	B Mandal	B P Nayak		Swetapadma Mishra/ Manoj Behera	

Subject(Theory)	Teacher(s)
English Communication Skill(ECS)	S Tripathy
Physics(Phy)	B P Nayak
Chemistry(Chem)	B Mandal
Mathematics(Math)	A K Panigrahi/Suchitra
Basics of Computer Engineering(BC)	A Panda
Robotics	A Nayak

Lab Class	Teacher(s)
English Com	S Tripathy/ P.Mishra
Computer L	A Panda
Phy Lab	B P Nayak/ R R Padhi
Chem Lab	B Mandal/ P Pandya
Robotics	Swetapadma Mishra/ Manoj Behera

HOBBY: Prof S. B. Pati to look after



Bridge Course Cumulative Attendance Sheet

Session	2021-22		Name of the Subject : Chemistry																	Bapaditya Mandal			
Year	1st		Group :		A															Name of the Faculty:		Bapaditya Mandal	
Sl NO	NAME OF THE STUDENT	Date & no of class																				Total	Remark
		1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20		
		24-05	25-05	27-05	01-06	02-06	04-06	08-06	09-06	11-06	18-06	20-06	21-06	24-06	25-06	27-06	30-06	02-07	03-07	06-07	07-07		
1	ARUNDHATI DAS	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20		
2	SATYANJIBI PRUSTY	A	A	A	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	16	
3	SHREEHARSH KESHAV	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
4	SIDHARTH SANKAR BEHERA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
5	SONALI PARIDA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
6	SOURAV KUMAR SHIKHAR	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
7	SUBHAM GOUDA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
8	TAPAS RANJAN PATRA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
9	TAPASWINEE DAS	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	19	
10	ANUBHAB RAY	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
11	PRITAM SEKHAR	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
12	RAKESH DAS	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
13	SATYABRATA DAS	P	P	P	P	P	P	P	A	A	P	P	P	P	P	P	P	P	P	P	P	18	
14	SAURAV PRASAD SWAIN	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
15	SIDHARTHA SAHOO	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
16	VIBHUTI BHUSHAN JENA	A	A	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	17	
17	VINAY KUMAR CHOUDHARY	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
18	ABHIPSA JENA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
19	ABHISHEK PRADHAN	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
20	ABHISHEK BARIK	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	18	
21	SUMAN KUMAR MEHER	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
22	SUMEET KUMAR NAIK	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
23	SUNIL MOHAPATRA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
24	SUPRAKASH GHOSH	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
25	SWATI KUMARI	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
26	SWAYANSIDDHA DASH	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
27	TAPAS KUMAR NAYAK	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
28	VYBRAJ MARNDI	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
29	ADARSH RAJULA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
30	PUJA SINGH	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
31	RASUDEV MALIK	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	19	
32	ANWESHA DAS	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	18	

33	DEBASISH BISWAL	P	P	P	D	P	P	P	P	D	D	P	D	D	D	P	D	P	D	P	P	20
34	DEBASISH RANSINGH	P	P	P	D	P	P	P	P	D	D	P	D	D	D	P	D	P	D	P	P	20
35	GAZALA PARWEEN	P	P	P	P	P	P	P	P	D	D	P	D	D	D	P	D	P	P	P	P	20
36	BHARAT CHANDRA GHUSI	P	P	P	D	P	P	P	P	D	D	P	D	D	D	P	D	P	P	P	P	20
37	DHIRENDRA NARAYAN SINGH DEO	P	P	P	D	P	P	P	P	D	D	P	D	D	D	P	D	P	P	P	P	20
38	SUBHASHREE SAHOO	P	P	A	P	P	P	P	P	D	D	P	D	D	D	P	P	A	D	D	P	18
39	SURHASIS MOHAPATRA	P	P	P	P	P	P	A	P	P	D	P	D	D	D	P	P	D	D	D	A	18
40	SURHASISH PATTANAİK	P	P	P	P	P	P	P	P	D	D	P	P	P	P	P	P	P	P	P	P	20
Signature of the Faculty		<i>P. Biswal</i>	<i>P. Ransingh</i>	<i>P. Gazala</i>	<i>P. Bharat</i>	<i>P. Dharendra</i>	<i>P. Subhashree</i>	<i>P. Surhasis</i>	<i>P. Surhasish</i>	<i>P. Biswal</i>	<i>P. Ransingh</i>	<i>P. Gazala</i>	<i>P. Bharat</i>	<i>P. Dharendra</i>	<i>P. Subhashree</i>	<i>P. Surhasis</i>	<i>P. Surhasish</i>	<i>P. Biswal</i>	<i>P. Ransingh</i>	<i>P. Gazala</i>	<i>P. Bharat</i>	<i>P. Dharendra</i>

[Handwritten Signature]

Signature of Coordinator / SPOC / HoD



Gandhi Institute For Technology (GIFT), Bhubaneswar

Bridge Course Cumulative Attendance Sheet

Session	2021-22	Name of the Subject :		Mathematics																	REMARKS	
Year	1st	Group :	B	Name of the Faculty													Parshuram Sahoo					
SL NO	NAME OF THE STUDENT	Date & no of class																				
		1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	Total
		22/05	24/05	28/05	29/05	01/06	04/06	06/06	09/06	11/06	21/06	23/06	25/06	29/06	01/07	03/07	05/07	09/07	10/07	14/07	17/07	
1	AKANKSHA SHARMA	P	a	P	P	P	P	P	P	P	P	P	P	P	a	a	a	P	P	P	P	16
2	ALIVA BHUNIA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
3	AMAN KUMAR SINGH	P	P	P	P	P	a	P	P	P	P	P	a	P	P	P	P	P	P	P	P	18
4	AMIT KUMAR JENA	P	P	P	P	a	a	a	a	a	P	P	P	P	P	P	P	P	P	P	P	15
5	ANURAG PRASAD	a	a	a	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	17
6	ARUN KUMAR	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
7	ARYA DEO MEHTA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
8	ASHISH KUMAR JAIN	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
9	ASHMIN KUMAR SAMAL	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
10	ASWINI KUMAR PATRA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
11	BHAGIRATHI PRADHAN	P	a	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	18
12	BHARTI PRASAD	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
13	BHAWANA KUMARI	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
14	BISMURTI BIODHIBRATA PATTAJOSHI	P	P	P	a	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	19
15	BJAJESH KUMAR	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	19
16	DEBASHREE KAR	P	P	P	P	P	P	P	a	P	P	P	P	P	P	P	P	P	P	P	P	19
17	DEEPAK KUMAR	P	P	P	P	P	P	P	P	P	P	P	a	P	P	P	P	P	P	P	P	19
18	DEWKEE RANI SAMAL	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
19	GAURAV UPADHYAY	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
20	GAUTAM SINGH	a	P	P	P	P	P	P	P	P	P	a	P	P	P	P	P	P	P	P	P	18
21	DHANESWAR GAHAN	P	P	P	P	P	P	a	a	P	P	P	P	P	P	P	P	P	P	P	P	18
22	MAHESH KUMAR BARIK	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
23	MD SHABBIR ALI	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
24	MD SHARIQ HASAN	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
25	NABIN KUMAR NAYAK	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
26	PHABHAKAR MALI	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
27	PRADDEEP KUMAR PARIDA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
28	PRAMOD KUMAR JENA	P	P	P	P	P	P	P	P	a	P	P	P	P	P	P	P	P	P	P	P	19
29	DEBASHREE PRADHAN	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
30	RADHANATH MOHANTY	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
31	RESAB SAHU	P	P	P	P	P	P	P	a	a	a	P	P	P	P	P	P	P	P	P	P	20
32	RISHI KUMAR	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	17
33	RISHALIN BEHERA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20
34	RISHAN LAL BASANTIA	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20



Gandhi Institute For Technology (GIFT), Bhubaneswar

Bridge Course Cumulative Attendance Sheet

Session	2021-22	Name of the Subject :		Mathematics																			
Year	1st	Group :	B	Name of the Faculty												Parshuram Sahoo							
SL NO	NAME OF THE STUDENT	Date & no of class																				REMARKS	
		1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20		Total
22	PRAS	22/05	24/05	28/05	01/06	04/06	06/06	09/06	11/06	21/06	23/06	25/06	29/06	01/07	03/07	05/07	09/07	10/07	14/07	17/07	18/07	17	
45	RISHIJEAN	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	17	
16	SANJIT KARAN	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
17	SANTANU KOLITRAY	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
48	SINHA LITAS	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
49	DEBANSU NAYAK	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
40	DEBANKAR ANAND	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	20	
Signature of the Faculty		PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	PR	

Pras

Signature of Coordinator / SPOC / HoD

Strategies adopted for slow learners

The institution takes care of the slow learners by adopting various strategies like remedial classes and doubt clearing classes. The students appear an entry level test called snap test for their classification into various categories. Based on the results of the test remedial and doubt clearing classes are arranged. The weak learners are provided with simple and standard lecture notes for better results. Specific faculties are assigned for counseling the students to monitor their academic performance regularly and interact frequently to understand, assist and support the students towards their better learning. Mentors communicate regularly with the parents and also send them SMS along with the progress of the student in regular intervals.



GANDHI INSTITUTE FOR TECHNOLOGY, BBSR

Ref No: GIFT/1st/09/2021

Date: 27th sep,2021

Notice for Snap Test(1st Semester)

This to inform you all that the **Snap Test** for the first year students will be conducted on **Monday (Dt: 27-10-2021)** according to following time schedule.

So you are requested collect the questions (Three sets) from the Dept. of BSH before time and to conduct the test in your respective class according to the schedule attached. At the end of the Test you are required to collect the answered question separately (as Physics, Chemistry and Mathematics sets) for ease in the evaluation process.

Timing for SNAP Test-20 21(Dt:27-10-21) Monday			
House	Room No	Faculty Name	Time
Anand	426	Prof. B. K. Rout	8.30am-9.30am
Kalam	427	Prof. Minati Pradhan	8.30am-9.30am
Behera	428	Prof. B. Mandal	8.30am-9.30am
Khosla	429	Prof. B.P.Nayak	8.30am-9.30am
Bhaba	433	Prof. Snigdha Tripathy	8.30am-9.30am
Viswesh	431	Prof. Minati Pradhan	11.30am-12.30pm

About The Test:

The test will be conducted in Physics, Chemistry and Mathematics subjects containing 20 multiple choice type questions in each subject for a total of (20+20+20=) 60 marks. Duration of the test is 45min.

Dr. ...
1st year Coordinator
(GIFT)
Bhubaneswar



DEPARTMENT OF BASIC SCIENCE AND HUMANITIES, GIFT, BBSR.

SNAP TEST · 2021 (1ST YEAR B.TECH STUDENTS)

DATE: FULL MARKS: 25 TIME: 30 minute.

NAME: ROLL NO. House:

MARKS OBTAINED:

SUBJECT - PHYSICS

TICK THE CORRECT ANSWER:(1x15=15)

Q 1. The dot product of two vectors is zero. What is the angle between them?

- (a) 45^0 (b) 60^0 (c) 90^0 (d) 0^0

Q.2. $\vec{A} = 3\hat{i} + 4\hat{j} + 5\hat{k}$. Find the magnitude of \vec{A} .

- (a) $6\sqrt{2}$ (b) $5\sqrt{2}$ (c) $7\sqrt{2}$ (d) none of these

Q.3. A charged particle of mass 'm' and the charge 'q' is projected with a velocity \vec{v} into uniform field of induction \vec{B} . The force experienced by the charged particle is

- (a) $\vec{F} = q(\vec{v} \times \vec{B})$ (b) $\vec{F} = q(\vec{v} \cdot \vec{B})$ (c) $\vec{F} = q/(\vec{v} \times \vec{B})$ (d) $\vec{F} = \vec{q} \times \vec{B}$

Q.4. If $\vec{A} \times \vec{B} = \vec{C}$, then

- (a) \vec{C} is perpendicular to \vec{A} only (b) \vec{C} is parallel to \vec{A}
(c) \vec{C} is perpendicular to both \vec{A} and \vec{B} (d) \vec{C} is parallel to both \vec{A} and \vec{B}

Q.5. Specific heat of a body depends upon :

- (a) mass of body (b) rise of temperature (c) amount of heat supplied (d) All of these

Q.6. A particle vibrating in S.H.M., while passing through the mean position will have :

- (a) maximum P.E. (b) maximum K.E. (c) both K.E. and P.E. maximum (d) none of these

Q.7. Angle between two unlike parallel forces is

- (a) 0 (b) $\pi/2$ (c) π (d) $\pi/3$

Q.8. A particle of mass 'm' has a momentum 'p'. Its kinetic energy will be

- (a) mp (b) $2p^2/m$ (c) p^2/m (d) $p^2/2m$

Q.9 The value of 'g' (acceleration due to gravity) is zero at

- (a) pole (b) equator (c) the centre of earth (d) none of these

Q.10. Numerical value of ϵ_0 (in S.I. unit) is

- (a) 8.85×10^{-10} (b) 9.5×10^{-10} (c) 8.85×10^{-12} (d) 9.5×10^{-12}

Q.11. If 'B' and 'H' are the magnetic induction and magnetic field then:

- (a) $H = \mu/B$ (b) $B = \mu/H$ (c) $H = B/\mu$ (d) $B = H/\mu$

Q.12. The magnitude of induced e.m.f. depends upon:

- (a) change in magnetic flux (b) rate of change of magnetic flux
(c) rate of change of electric field (d) none of these

Q.13. If the focal length of the lens is 10cm, the power of lens is

- (a) 1D (b) 0.1D (c) 0.01D (d) 10D

Q.14. Condition for constructive interference is that the path difference of rays from two sources should be :

- (a) even multiple of $\lambda/2$ (b) odd multiple of $\lambda/2$
(c) an integral multiple of $\lambda/2$ (d) none of these

Q.15. If 'V' is the potential difference, 'I' is the current and 'R' is the resistance, then which relation is correct?

- (a) $R = VI$ (b) $R = V/I$ (c) $V = I/R$ (d) none of the

Q.16. what is Biot-Savart's law? Derive its expression in S.I. unit. Express it in vector form. (5marks)

Q.17. State Newton's three laws of motion and prove that Force = mass x acceleration from (5marks)



DEPARTMENT OF BSH, GIFT, BBSR
SNAP TEST FOR 1ST YEAR B.TECH (2021-22)

NAME OF THE STUDENT:-

DURATION:- 30 Minutes

ROLL NO:-

HOUSE:

FULL MARKS: 25

MARKS OBTAINED:

Sub: CHEMISTRY

1. The HCP arrangement is described by :
(i) AB AB..... (ii) ABC ABC..... (iii) AB A BA..... (iv) ABC CBA...
2. The half life of zero order reaction is
(i) Independent of initial concentration
(ii) Directly proportional to square of concentration
(iii) Inversely proportional to square of concentration
(iv) Directly proportional to the initial concentration
3. The d-block elements are known as:
(i) Alkali (ii) Transition (iii) Alkaline earth (iv) Noble
4. Which of the following is a metalloid:
(i) Pb (ii) Br (iii) Sb (iv) Al
5. Electro negativity refers to the tendency of an atom to
(i) Lose electrons
(ii) Attract electrons involved in chemical bonding
(iii) Repel electrons
(iv) Share electrons with other atoms by covalent bonding
6. The d-block elements are known as:
(i) Alkali (ii) Transition (iii) Alkaline earth (iv) Noble
7. Which of the following is a metalloid:
(i) Pb (ii) Br (iii) Sb (iv) Al
8. Electro negativity refers to the tendency of an atom to
(i) Lose electrons
(ii) Attract electrons involved in chemical bonding
(iii) Repel electrons
(iv) Share electrons with other atoms by covalent bonding
9. Water is a liquid because of
(i) Covalent bond (ii) intramolecular hydrogen bond
(iii) Intermolecular hydrogen bond (iv) Ionic bond
10. Milk is a
(i) Gel (ii) Solution (iii) emulsion (iv) none of the above
11. Evaporation of water is

- (i) An endothermic change
- (ii) An exothermic change
- (iii) A process where no heat change occurs
- (iv) A process accompanied by chemical reaction.

12. What is the oxidation number of sulphur in H_2SO_4 ?

- (i) +2 (ii) -2 (iii) +6 (iv) -4

13. The heat change at constant pressure q_p is equal to

- (i) ΔU (ii) ΔG (iii) RT (iv) ΔH

14. Which molecule is paramagnetic?

- (i) O_2 (ii) N_2 (iii) O_2^- (iv) F_2

15. Different crystalline forms of the same elements are called

- (i) Isomers (ii) Isotopes (iii) Allotropes (iv) Isotones

LONG QUESTIONS

1. Explain briefly Hess's law of constant heat summation.

2. Distinguish between order and molecularity of a reaction.



GANDHI INSTITUTE FOR TECHNOLOGY, BBSR

Ref No: GIFT/1st/02/: 2022

Date: 21 February, 2022

Notice for Remedial Classes (1st Year)

It is hereby informed to all 1st year B.Tech students that Remedial Classes for the subjects Math-II, PL, BE/BEE, BME/BCE & Physics/Chemistry shall start from 2nd March, 2020 (Saturday) from 4.30pm to 6.30pm. The details of the group list & time table is available on the notice board.

Any students whose name is enlisted in the list, but not interested to attend the classes may give their submission in writing through the proctor concerned. Similarly any student whose name does not find a place in the list, but wants to attend such classes may also submit the request in writing through the proctor.

All students are advised to make full use of the classes as personal help and doubt clearance shall be provided by the faculties in these classes.

Dr. ...
1st year Coordinator

Copy to: All Notice Boards/Hostel Notice Boards/ By mail to Principal / All Deans/Vice Chairman / all Proctors & faculties of 1st year.



Gandhi Institute For Technology (GIFT), Bhubaneswar

Remedial Class Attendance sheet

Sub-Chemistry

Sl. No	Regn.No	Name	Year	04-Mar	07-Mar	11-Mar	15-Mar	19-Mar	23-Mar	27-Mar	31-Mar	05-Apr	Total Attendance
1	1901298220	SUCHISMITA BEBARTA	1st	P	P	P	A	P	P	P	P	P	7
2	1901298289	ANKIT SAHU	1st	A	P	P	P	P	A	P	P	P	7
3	1901298290	ARPITA BEHERA	1st	P	P	P	P	A	P	P	P	A	7
4	1901298291	ARUP KUMAR JENA	1st	P	A	P	P	P	P	P	P	P	7
5	1901298292	ASHIS ASUTOSH NAYAK	1st	P	P	P	P	P	P	P	P	P	7
6	1901298293	BISHNU NARAYAN MOHAPATRA	1st	P	P	P	P	P	P	P	P	P	7
7	1901298294	DEBANSHU SHEKHAR BISWAL	1st	P	P	P	P	P	P	P	P	P	7
8	1901298295	MAHENDRA PRATAP SAHOO	1st	P	P	P	A	P	P	P	P	P	7
9	1901298296	MANOJ PRADHAN	1st	P	P	P	A	P	P	P	P	P	7
10	1901298297	NETAI LOHAR	1st	A	P	P	A	P	P	P	P	P	7
11	1901298298	NRUPENDRA ROUT	1st	P	P	A	P	P	P	P	P	P	7
12	1901298299	PREM KUMAR	1st	P	P	P	P	P	P	P	P	P	7
13	1901298300	RANI KUMARI	1st	A	P	P	P	P	P	P	A	P	7
14	1901298301	SATYA PRAKASH SAHOO	1st	P	P	P	P	P	P	P	P	P	7
15	1901298302	SONALI BEHERA	1st	P	P	P	P	P	P	P	P	A	7
16	1901298303	SOUFYADARSAN ACHARYA	1st	P	P	P	A	P	P	P	P	P	7
17	1901298304	SOURAV MILLAN ROUT	1st	P	P	P	P	P	P	P	P	P	7
18	1901298305	SUCHISMANTA BARIK	1st	P	P	P	P	P	P	P	P	P	7
19	1901298306	SUMAN HATI	1st	P	P	A	P	P	P	P	A	P	7
20	1901298307	SUSHREE SANGITA BEHERA	1st	P	P	P	P	P	P	P	P	P	7
21	1901298308	SUSRATA DAS	1st	A	P	P	A	P	P	P	P	P	7
Signature of the faculty				R Paul	R Paul	R Paul	R Paul	R Paul	R Paul	R Paul	R Paul	R Paul	R Paul

Signature of the co-ordinator



Gandhi Institute For Technology (GIFT), Bhubaneswar

Remedial Class attendance sheet

													Sub-PL
Sl No	Regn.No	Name	Year	4 Mar	7 Mar	11 Mar	15 Mar	19 Mar	23 Mar	27 Mar	31 Mar	5 Apr	Total Attendance
	1901298220	SUCHISMITA BEBARTA	1st	a	P	P	P	a	P	P	P	P	7
1	1901298289	ANKIT SAHU	1st	P	P	P	P	P	P	P	P	P	9
2	1901298290	ARPITA BEHERA	1st	P	P	P	P	P	a	P	P	P	8
3	1901298291	ARUP KUMAR JENA	1st	a	a	a	P	P	P	P	P	P	9
4	1901298292	ASHIS ASUTOSH NAYAK	1st	P	P	P	P	P	P	P	P	P	9
5	1901298293	BISHNU NARAYAN MOHAPATRA	1st	P	P	a	P	P	a	P	P	P	8
6	1901298294	DEBANSHU SHEKHAR BISWAL	1st	P	P	P	P	P	a	P	P	P	8
7	1901298295	MAHENDRA PRATAP SAHOO	1st	P	P	P	P	P	a	P	P	P	8
8	1901298296	MANOJ PRADHAN	1st	P	P	P	P	P	P	P	P	P	9
9	1901298297	NETAI LOHAR	1st	P	P	P	P	P	P	P	P	P	9
10	1901298298	NRUPENDRA ROUT	1st	P	P	a	P	P	P	P	P	P	8
11	1901298299	PREM KUMAR	1st	P	P	P	P	P	P	P	P	P	9
12	1901298300	RANI KUMARI	1st	P	P	P	P	P	P	P	P	P	9
13	1901298301	SATYA PRAKASH SAHOO	1st	P	P	P	P	P	P	P	P	P	9
14	1901298302	SONALI BEHERA	1st	P	P	P	P	P	P	P	P	a	8
15	1901298303	SOUFYADARSAN ACHARYA	1st	P	P	P	P	P	P	P	P	P	9
16	1901298304	SOURAV MILLAN ROUT	1st	P	P	P	P	P	P	P	P	P	9
17	1901298305	SUCHISMANTA BARIK	1st	a	P	P	P	P	P	P	P	P	9
18	1901298306	SUMAN HATI	1st	P	P	P	P	P	P	P	P	P	9
19	1901298307	SUSHREE SANGITA BEHERA	1st	P	P	P	P	P	P	P	P	P	9
20	1901298308	SUSRATA DAS	1st	P	P	P	P	P	P	P	P	P	9
Signature of the faculty				<i>Exipathy</i>	<i>Exipathy</i>	<i>Exipathy</i>	<i>Exipathy</i>	<i>Exipathy</i>	<i>Exipathy</i>	<i>Exipathy</i>	<i>Exipathy</i>	<i>Exipathy</i>	

Roal -
Signature of the co-ordinator

STUDY MATERIAL

ON

BE

(B-TECH 1st/2nd semester)

MODULE- I,II,III,IV



PREPARED BY:

Prof. SAUMENDRA KU. BEHERA
Prof. SUBRAT KU. PANDA

DEPARTMENT OF ECE.

**GANDHI INSTITUTE FOR TECHNOLOGY
BHUBANESWAR**

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The atoms are the basic building blocks of all matter. By based on atoms arrangement, Niels Bohr put forward his model i.e. called Bohr's Atomic model.

Bohr's Atomic model \Rightarrow

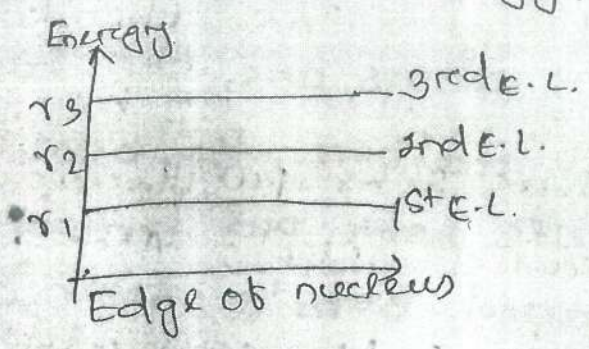
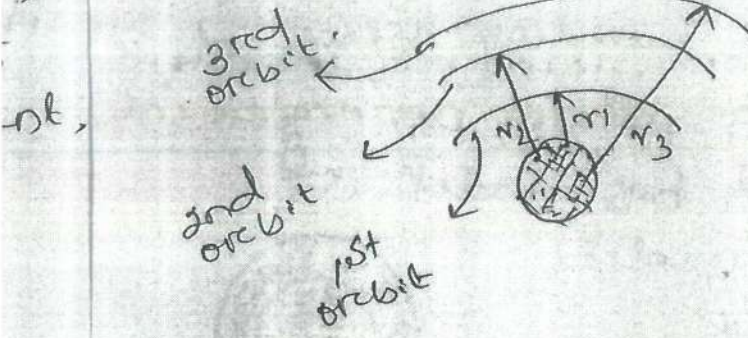


- An Atom comprises of ~~truly~~ positively charged nucleus surrounded which ~~truly~~ negatively charged electrons revolve in various ^{circular} orbitals.
- Nucleus is central part of an atom & contains protons & neutrons.
- The electrons in each allowed orbit have fixed amount of energy. The higher the orbit, greater is energy of electrons.
- If additional energy is supplied to electron, it is jumped to higher orbit. i.e. called excitation state.
- This state doesn't last long, becuz electron drops down the original lower orbit. As it drops, it gives back acquired energy in form of heat, light or other radiation.
- Electrons can revolve only in allowed orbit (K, L, M...)

Energy Levels :-

Each orbit has fixed amount of energy associated with it. The electrons revolving 22

Higher the orbit, greater is its energy.



The 1st orbit indicates, 1st energy level, 2nd orbit indicates the 2nd energy level.

Single isolated atom, the energy is represented by horizontal line i.e. called energy level.

Energy bands :-

- Energy level terms associated with a system of electrons in one isolated atom.
- But in a solid, atoms are greatly influenced by closely packed neighbouring atoms.
- The closeness of atoms results in intermixing of electron of neighbouring atoms. Due to intermixing, energy level increases.
- Under these conditions, energy levels that may be occupied by electrons merge into bands of energy levels.

Range of energies possessed by electrons of same orbit in a solid is called energy band.

